

JEE Main 5th April Session 2 - Memory Based Paper

5th April Session 2



Questions

Q.1. Which of the following is not correct about stopping potential?

- A) It depends on the frequency of the incident light.
- B) It depends on the nature of the material.
- C) It is equal to $\frac{K.E.}{e}$.
- D) It increases with intensity of the incident light.

Answer: It increases with intensity of the incident light.

Solution: The expression for the stopping potential of a metal can be written as

$$eV_s = h\nu - \phi \dots(1)$$

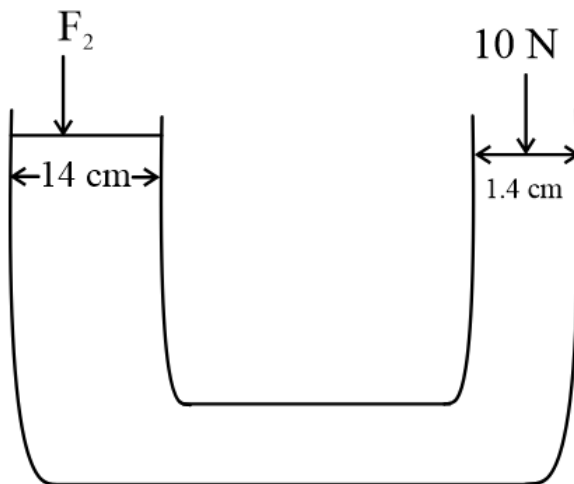
From the above equation, it can be concluded that the stopping potential depends on the nature of the material and the frequency of the incident light.

Also, the kinetic energy of the electron is given by

$$K.E. = eV_s$$
$$\Rightarrow V_s = \frac{K.E.}{e}$$

The stopping potential does not depend on the intensity of the light. Hence, this is the correct option.

Q.2. Find the value of F_2 in newton unit for the following diagram.



- A) 100
- B) 10
- C) 1000
- D) 10000

Answer: 1000



Solution: For the left arm of the tube, from Pascal's law, it follows that

$$P = \frac{F_2}{\pi(14 \times 10^{-2})^2}$$
$$= \frac{4F_2}{196 \times 10^{-4}\pi} \dots (1)$$

For the right arm of the tube, it follows that

$$P = \frac{10 \text{ N}}{\pi(1.4 \times 10^{-2})^2}$$
$$= \frac{40 \text{ N}}{1.96 \times 10^{-4}\pi} \dots (2)$$

From equations (1) and (2),

$$\frac{4F_2}{196\pi} = \frac{40 \text{ N}}{1.96\pi} \Rightarrow F_2 = 1000 \text{ N}$$

Q.3. Given that the current varies with time through an inductor as $I = (3t + 8) \text{ A}$. If the inductor is connected across a source of emf $\varepsilon = 12 \text{ V}$, find the self inductance (in H) of the inductor.

A) 2

B) 12

C) 4

D) 14

Answer: 4

Solution: Given,

$$I = 3t + 8 \dots (1)$$

The formula to calculate the self-inductance of the inductor is given by

$$\varepsilon = L \frac{dI}{dt} \dots (2)$$

From equations (1) and (2), it follows that

$$12 = L \times \frac{d}{dt}(3t + 8)$$
$$\Rightarrow L = \frac{12}{3}$$
$$= 4 \text{ H}$$

Q.4. The maximum height reached by a projectile is 64 m. If the initial velocity is halved, then calculate the new maximum height of the projectile in meters.

A) 12

B) 10

C) 16

D) 20

Answer: 16



Solution: The formula to calculate the maximum height attained by the projectile is

$$H = \frac{v^2 \sin^2 \theta}{2g} \dots (1)$$

where, v is the initial velocity and θ is the angle of projection.

Substitute the new value of velocity into equation (1) to obtain the new value (H') of maximum height.

$$\begin{aligned} H' &= \frac{\left(\frac{v}{2}\right)^2 \sin^2 \theta}{2g} \\ &= \frac{1}{4} \frac{v^2 \sin^2 \theta}{2g} \\ &= \frac{H}{4} \\ &= \frac{64 \text{ m}}{4} \\ &= 16 \text{ m} \end{aligned}$$

Q.5. The Van der Waal's gas equation is expressed as $\left(P - \frac{a}{V^2}\right)(V - b) = nRT$, where symbols have their usual meaning. Then dimension of $\frac{a}{b^2}$ is

- A) $[ML^{-1}T^{-2}]$ B) $[M^{-1}LT^{-2}]$
 C) $[ML^{-2}T^{-2}]$ D) $[ML^{-1}T^{-1}]$

Answer: $[ML^{-1}T^{-2}]$

Solution: Only similar nature quantities can be added or subtracted, so

It means V and b must have same dimensions, so

$$[b] = [V] = [L^3]$$

Similarly p and $\frac{a}{V^2}$ have same dimensions, so

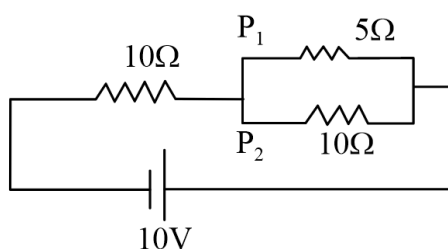
$$\left[p \right] = \frac{[a]}{[V]^2} = [ML^{-1}T^{-2}]$$

$$[a] = [ML^5T^{-2}]$$

Hence,

$$\begin{aligned} \left[\frac{a}{b^2} \right] &= \frac{[ML^5T^{-2}]}{[L^6]} \\ &= [ML^{-1}T^{-2}] \end{aligned}$$

Q.6. The ratio of heat dissipated $\left(\frac{P_1}{P_2}\right)$ per second through 5Ω and 10Ω is given as



- A) 1 : 2 B) 1 : 1
 C) 2 : 1 D) 4 : 1



Answer: 2 : 1

Solution: Since both branches are in parallel, therefore potential drop will be equal.

So for power, we can write

$$\frac{P_1}{P_2} = \frac{\left(\frac{V^2}{R_1}\right)}{\left(\frac{V^2}{R_2}\right)} = \frac{R_2}{R_1} = \frac{10}{5} = 2 : 1$$

Q.7. During an adiabatic process, if pressure of gas is found to be proportional to cube of its absolute temperature, then $\frac{C_p}{C_v}$ for gas is

- A) $\frac{3}{2}$ B) $\frac{5}{3}$
C) $\frac{7}{5}$ D) $\frac{9}{7}$

Answer: $\frac{3}{2}$

Solution: Given:

$$\begin{aligned}P &\propto T^3 \\ \Rightarrow P &\propto (PV)^3 \\ \Rightarrow P^2 V^3 &= \text{constant} \\ \Rightarrow P V^{\frac{3}{2}} &= \text{constant}\end{aligned}$$

Now, comparing the above equation with the standard equation ($PV^\gamma = \text{constant}$) for adiabatic process, we get $\gamma = \frac{3}{2} = \frac{C_p}{C_v}$.

Q.8. A block of mass 50 kg is moving with speed of 10 m s⁻¹ on a rough horizontal surface (friction coefficient = 0.3). Find the kinetic friction acting on the object.

- A) 16 N B) 150 N
C) 167 N D) 500 N

Answer: 150 N

Solution: Normal force acting on the object will be $N = mg$.

Now, kinetic friction is given by the expression, $f = \mu N = \mu mg = 0.3 \times 50 \times 10 = 150 \text{ N}$.

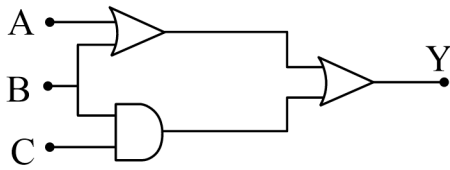
Q.9. A truck is moving from rest with constant power P . If the displacement of the truck is proportional to t^n , where is t is time, find n .

- A) $\frac{1}{2}$ B) 2
C) $\frac{3}{2}$
D) $\frac{5}{2}$

Answer: $\frac{3}{2}$



Q.12. For what Boolean values of A, B & C, the given logic gate gives an output of zero.



- A) $A = 1, B = 0, C = 1$ B) $A = 0, B = 1, C = 1$
- C) $A = 1, B = 1, C = 1$ D) $A = 0, B = 0, C = 1$

Answer: $A = 0, B = 0, C = 1$

Solution: The output (X) of the first OR gate is given by

$$X = A + B$$

The output (X') of the AND gate is given by

$$X' = B \cdot C$$

Hence, the net output (Y) is given by

$$\begin{aligned} Y &= X + X' \\ &= A + B + B \cdot C \\ &= A + B(1 + C) \\ &= A + B \end{aligned}$$

From the above calculation, it is clear that the value of the final output does not depend C. So, for the output to be zero, the values of A and B have to be zero.

Hence, this is the correct option.

Q.13. In sonometer, fundamental frequency changes from 400 Hz to 500 Hz keeping the same tension. Find the percentage change in length.

- A) 10% B) 5%
- C) 20% D) 40%

Answer: 20%

Solution: For the old fundamental frequency,

$$\begin{aligned} F &= \frac{v}{2l_1} \\ \Rightarrow l_1 &= \frac{v}{2F} \\ &= \frac{v}{800} \dots (1) \end{aligned}$$

For the new fundamental frequency,

$$\begin{aligned} F' &= \frac{v}{2l_2} \\ \Rightarrow l_2 &= \frac{v}{1000} \dots (2) \end{aligned}$$

Hence, the required percentage error (p) in length can be calculated as follows:

$$\begin{aligned} |p| &= \left| \frac{l_2 - l_1}{l_1} \times 100\% \right| \\ &= \left| \frac{\frac{v}{1000} - \frac{v}{800}}{\frac{v}{800}} \times 100\% \right| \\ &= \left[1 - \frac{8}{10} \right] \times 100\% \\ &= 20\% \end{aligned}$$

Q.14. The Electrostatic force (\vec{F}_1) and magnetic force (\vec{F}_2) acting on charge q moving with a velocity \vec{v} in an electric field \vec{E} and magnetic field \vec{B} is given by



- A) $\vec{F}_1 = \vec{v} \times \vec{E}$, $\vec{F}_2 = q(\vec{B} \times \vec{v})$ B) $\vec{F}_1 = q\vec{B}$, $\vec{F}_2 = q(\vec{E} \times \vec{v})$
- C) $\vec{F}_1 = q\vec{E}$, $\vec{F}_2 = q(\vec{v} \times \vec{B})$ D) $\vec{F}_1 = q\vec{E}$, $\vec{F}_2 = q(\vec{B} \times \vec{v})$

Answer: $\vec{F}_1 = q\vec{E}$, $\vec{F}_2 = q(\vec{v} \times \vec{B})$

Solution: The force due to the electric field on a charge is built into its definition, $\vec{F}_1 = q\vec{E}$.

Magnetic fields can exert a force on an electric charge only if it moves and the force is given by, $\vec{F}_2 = q(\vec{v} \times \vec{B})$.

Q.15. 20 Ω resistance is divided into 10 equal parts and they are connected in parallel. The overall resistance (in Ω) is

- A) 0.2 B) 0.8
- C) 0.4 D) 0.6

Answer: 0.2

Solution: When the 20 Ω resistor is divided into ten parts, each resistance becomes 2 Ω .

Hence, the equivalent resistance (R) can be calculated as follows:

$$\begin{aligned} \frac{1}{R} &= \frac{1}{2} + \frac{1}{2} + \dots + 10 \text{ terms} \\ &= \frac{10}{2} \\ \Rightarrow R &= \frac{2}{10} \\ &= 0.2 \Omega \end{aligned}$$

Q.16. For a series LCR circuit having AC source of 220 V, 50 Hz, it is observed that the voltage across the inductor which has self inductance of 10 mH is 31.4 V. Find the current in circuit.

- A) 10 mA B) 63 mA
- C) 10 A D) 63 A

Answer: 10 A

Solution: Voltage across the inductor,

$$\begin{aligned} V_L &= IX_L \\ \Rightarrow 31.4 &= I(\omega L) \\ \Rightarrow 31.4 &= I(2\pi \times 50 \times 10 \times 10^{-3}) \\ \Rightarrow I &= 10 \text{ A} \end{aligned}$$

Q.17. In the hydrogen spectrum, the wavelength of the shortest line from the Lyman series is 916 Å, then find the longest wavelength of the Balmer series.

- A) 65 Å B) 95 Å
- C) 6595 Å D) 9565 Å

Answer: 6595 Å



Solution: For the shortest wavelength in Lyman series, it can be written that

$$\frac{1}{\lambda_L} = R \left(\frac{1}{1^2} - \frac{1}{\infty} \right)$$

$$= R \dots (1)$$

For the longest wavelength in Balmer series, it follows that

$$\frac{1}{\lambda_B} = R \left(\frac{1}{2^2} - \frac{1}{3^2} \right)$$

$$= \frac{5R}{36} \dots (2)$$

Equations (1) and (2) imply that

$$\frac{1}{\lambda_B} = \frac{5}{36\lambda_L}$$

$$\Rightarrow \lambda_B = \frac{36}{5} \times 916 \text{ \AA}$$

$$\approx 6595 \text{ \AA}$$

Q.18. Find out E_{cell} of the given cell $M | M^{2+} || X^{2-} | X$

$$E^\circ_{M^{2+}|M} = 0.34 \text{ V}$$

$$E^\circ_{X|X^{2-}} = 0.46 \text{ V}$$

- A) 0.80 V B) 0.12 V
C) -0.12 V D) -0.80 V

Answer: 0.12 V

Solution: E_{cell} is the difference between standard reduction potential of cathode and standard reduction potential of anode.

$$E^\circ_{(X|X^{2-})} - E^\circ_{(M^{2+}|M)}$$

$$= 0.46 - 0.34$$

$$= 0.12 \text{ V}$$

Q.19. Angular momentum of an electron in an orbit of radius R of a hydrogen atom is directly proportional to

- A) $\frac{1}{R}$ B) R
C) $\frac{1}{\sqrt{R}}$
D) \sqrt{R}

Answer: \sqrt{R}

Solution: The angular momentum of Bohr's orbit is

$$L = mvr = \frac{nh}{2\pi}$$

The n^{th} orbit radius can be calculated as follows:

$$R = 0.529 \times \frac{n^2}{Z} \text{ \AA}$$

$$n \propto \sqrt{ZR}$$

$$\Rightarrow L \propto \sqrt{R}$$

Q.20. Which of the following is true regarding coagulation of the egg?

- A) Primary structure does not change B) Secondary structure does not change
C) Tertiary structure does not change D) Denaturation of protein does not occur

Answer: Primary structure does not change



Solution: Denaturation of proteins protein present in egg white has an unique three dimensional structure. When it is subjected to physical change like change in temperature. i.e., on boiling, coagulation of egg white occurs due to denaturation of protein. During denaturation hydrogen bonds are disturbed due to this globules unfold and helix gets uncoiled and protein loses its biological activity.

During denaturation, secondary and tertiary structures are destroyed but primary structure remains intact. The coagulation of egg white on boiling is a common example of denaturation.

Q.21. Find out value of $\frac{C_p}{C_v}$ for an ideal gas undergoing reversible adiabatic process for which $P \propto T^3$.

- A) 3/2 B) 5/4
 C) 4/3 D) 5/3

Answer: 3/2

Solution: For an adiabatic reversible process,

$$PV^\gamma = k(\text{constant})$$

$$\frac{C_p}{C_v} = \gamma \text{ for an ideal gas.}$$

From the ideal gas equation, $PV = nRT$

$$\text{Given that } \frac{P}{T^3} = \text{constant}$$

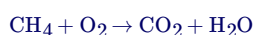
$$\Rightarrow \frac{P}{(PV)^3} = \text{constant}$$

$$\Rightarrow \frac{1}{P^2V^3} = \text{constant}$$

$$\Rightarrow PV^{3/2} = \text{constant}$$

$$\text{Hence, } \frac{C_p}{C_v} = 3/2$$

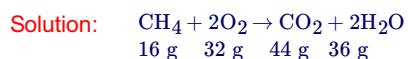
Q.22. For the reaction,



How many moles of methane is required for the formation of 11 g of CO_2 .

- A) 0.25 B) 0.45
 C) 0.32 D) 0.36

Answer: 0.25



From the above reaction, we can say that one mole of methane gives one mole of carbon dioxide.

$$\text{Number of mole of methane required for the formation of 11 g of } \text{CO}_2 = \frac{11}{44} = \frac{1}{4} = 0.25 \text{ moles}$$

Q.23. Assertion: NF_3 have less dipole moment than NH_3 .

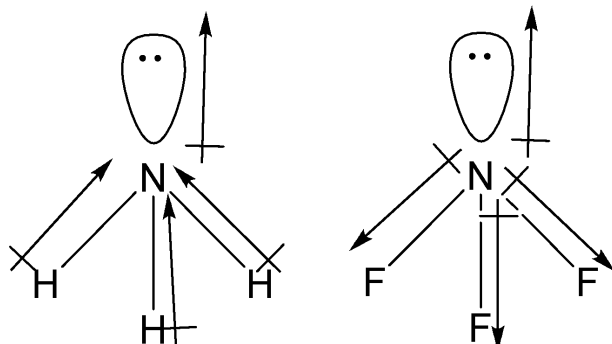
Reason: F is more electronegative and bond moment of N – H bond is in the same direction of lone pair.

- A) Both assertion and reason are correct and reason is the correct explanation for the assertion
 B) Both assertion and reason are correct and reason is not the correct explanation for the assertion
 C) Assertion is correct and reason is incorrect
 D) Assertion is incorrect and reason is correct

Answer: Both assertion and reason are correct and reason is the correct explanation for the assertion



Solution: In NH_3 , the dipole moment vector of the bond and the lone pairs are in the same direction. But in NF_3 , the dipole moment vector of lone pairs and bond pairs are opposite in direction. So, the net dipole moment will be the subtractive effect of the two. Hence, the dipole moment of NH_3 is larger than NF_3 .



Fluorine is more electronegative than nitrogen so all the electron will move towards fluorine. Due to this all of the vectors will cancel each other and the summation will come minimum.

Q.24. Given elementary reaction



The initial pressures of A and B respectively are 1.5 atm and 0.7 atm. After time t, the pressure of C is 0.5 atm. Find the ratio of initial rate and rate after time t.

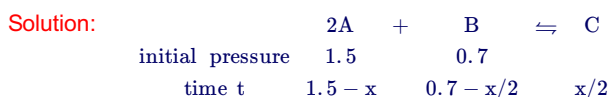
A) $\frac{63}{2}$

B) $\frac{33}{2}$

C) $\frac{23}{2}$

D) $\frac{31}{2}$

Answer: $\frac{63}{2}$



Given that
 $x/2 = 0.5 \text{ atm}$
 $\Rightarrow x = 1 \text{ atm}$

Now partial pressures of A & B after time t is 0.5 atm and 0.2 atm respectively.

Now, rate law can be written as

$$R = k[A]^2[B]$$

$$R_1 (\text{initial rate}) = k[1.5]^2[0.7]$$

$$R_2 (\text{rate at time t}) = k[0.5]^2[0.2]$$

$$\frac{R_1}{R_2} = \frac{[1.5]^2[0.7]}{[0.5]^2[0.2]} = 31.5$$

Q.25. Correct order of increasing atomic size of 13th group element will be:

A) $\text{Tl} > \text{In} > \text{Al} > \text{Ga} > \text{B}$

B) $\text{Tl} < \text{In} > \text{Al} > \text{Ga} > \text{B}$

C) $\text{Tl} > \text{In} < \text{Al} > \text{Ga} > \text{B}$

D) $\text{In} > \text{Tl} > \text{Ga} > \text{Al} > \text{B}$

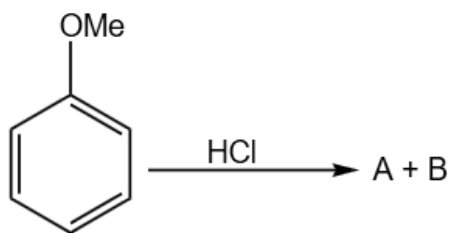
Answer: $\text{Tl} > \text{In} > \text{Al} > \text{Ga} > \text{B}$

Solution: Atomic and ionic radii of group 13 elements are lower than those of alkaline earth metals of group 2 primarily due to greater nuclear charge of group 13 elements as compared to group 2 elements. On moving down the group the atomic radius of Ga is slightly lower than that of Al. This is due to the presence of d - electrons in Ga which do not shield the nucleus effectively. As a result, the electrons in Ga experience greater force of attraction by the nucleus than in Al and hence the atomic radius of Ga 135 pm is slightly less than that of Al 143 pm. Thus, the increasing order of atomic radii of group 13 elements is $\text{B} < \text{Ga} < \text{Al} < \text{In} < \text{Tl}$.

Hence, the correct option is A



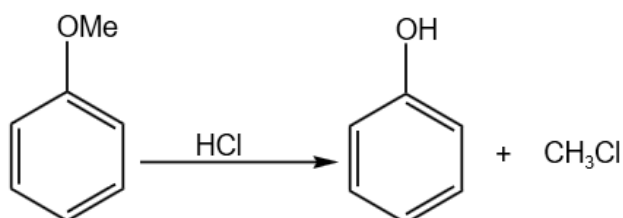
Q.26. Identify the products A and B in the following reaction.



- A) Phenol and methyl chloride B) Phenol and methanol
C) Chlorobenzene and methanol D) Chlorobenzene and methyl chloride

Answer: Phenol and methyl chloride

Solution: Anisole on reaction with hydrogen chloride gives phenol and methyl chloride as the major products. Phenol is major product because of double character between carbon of benzene ring and oxygen.



Q.27. Match Column I with column II

Column I	Column II
1. ICl	P. T-shape
2. ICl ₃	Q. Pentagonal bipyramidal
3. ClF ₅	R. Linear
4. IF ₇	S. Square pyramidal

- A) 1 – R, 2 – P, 3 – S, 4 – Q B) 1 – R, 2 – P, 3 – Q, 4 – S
C) 1 – P, 2 – R, 3 – S, 4 – Q D) 1 – R, 2 – S, 3 – P, 4 – Q

Answer: 1 – R, 2 – P, 3 – S, 4 – Q

Solution: ICl has linear shape due to lone of electrons.

The given compound is ICl₃. Due to the presence of two lone pairs on iodine, the molecular geometry of this compound is T-shaped. The electron pair geometry is trigonal bipyramidal.

ClF₅ If the ligand atoms were connected, the resulting shape would be that of a pyramid with a square base.

IF₇, or Iodine heptafluoride is an interhalogen compound. It has a coordination number of 7. It was predicted by the VSEPR theory that IF₇ has an unusual pentagonal bipyramidal structure as the central iodine atom undergoes sp³d³ hybridisation.

Q.28. From the given information, calculate the enthalpy of formation of 2 moles of C₆H₆(l) at 25°C.

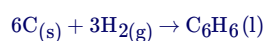
Given $\Delta_{\text{C}}\text{H}_{\text{C}_6\text{H}_6(\text{l})} = -3264.6 \text{ kJ/mol}$, $\Delta_{\text{C}}\text{H}_{\text{C}(\text{s})} = -393.5 \text{ kJ/mol}$, $\Delta_{\text{C}}\text{H}_{\text{H}_2\text{O}(\text{l})} = -285.83 \text{ kJ/mol}$

- A) 246.11 kJ/mol B) 46.11 kJ/mol
C) -124.5 kJ/mol D) 124.5 kJ/mol



Answer: 46.11 kJ/mol

Solution: The formation reaction of benzene is



The enthalpy of formation of benzene can be calculated as follows,

$$\Delta H_f(\text{C}_6\text{H}_6(l)) = 6 \times \Delta H_c(\text{C}_{(s)}) + 3 \times \Delta H_c(\text{H}_{2(g)}) - \Delta H_c(\text{C}_6\text{H}_6(l))$$

$$\Rightarrow \Delta H_f(\text{C}_6\text{H}_6(l)) = 6 \times (-393.5) + 3 \times (-285.83) - (-3264.6)$$

$$\Rightarrow \Delta H_f(\text{C}_6\text{H}_6(l)) = 46.11 \text{ kJ/mol}$$

Q.29. Which of the following is an acidic oxide?

A) N_2O_3 B) NO

C) CO

D) CaO

Answer: N_2O_3

Solution: An oxide that exhibits the properties of acids is called acidic oxide. Since they are acidic, they react with a base to form a salt. All the nitrogen oxides are acidic except nitric oxide and nitrous oxide, which are neutral oxides. Carbon monoxide is a neutral oxide and calcium oxide is a basic oxide.

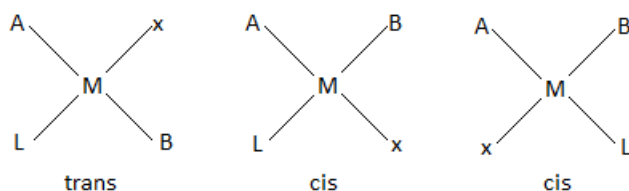
Q.30. If square planar complex $\text{M}(\text{ABXL})$ has all four unidentate ligands, then find out its total number of geometrical isomers.

A) 2 B) 3

C) 4 D) 0

Answer: 3

Solution: The square planar complex of the type MABXL (where A, B, X, L are unidentates) exhibit three geometrical isomers. Out of three isomers, two isomers are cis and one isomer is trans.



Q.31. How many of the following have zero dipole moment?

H_2S , CH_4 , NH_3 , BF_3 , SO_2 , NF_3

Answer: 2

Solution: The bond moments act from Hydrogen towards sulphur and the moment from sulphur acts towards the lone pair of electrons. Both the bond moments act in the same direction and thus the resultant moment adds up vectorially.

The dipole moment of methane is zero because the molecule is a symmetrical tetrahedron. Therefore, the dipole moments between carbon and hydrogen cancel out each other, leading to a net dipole moment of zero.

The NH_3 molecule has a dipole moment because the three N-H bond dipoles add to give a net dipole pointing upward to the NH_3 structure.

The fluorine is the most electronegative atom, so the movement of dipole is towards the fluorine atom. As the magnitude is the same for the two B-F bonds which are opposite to the third B-F bond, it will cancel the dipole of the third B-F bond and as a result the dipole will be zero.

The geometry of SO_2 is bent because the lone pair on the central sulphur atom pushes the two bonds down for a bond angle of due to the electrons repulsion. This is what makes SO_2 a polar molecule because the dipole moments of both bonds do not cancel out,

In case of NF_3 the dipole moment is not zero because of presence of a lone pair of electrons on Nitrogen atom.



Q.32. In an atom, how many electrons can have:

$$(i) n = 4 \quad (ii) m_l = 1 \quad (iii) m_s = \frac{1}{2}$$

Answer: 3

Solution: $4s, m_l = 0$
 $4p, m_l = -1, 0, +1$
 $4d, m_l = -2, -1, 0, +1, +2$
 $4f, m_l = -3, -2, -1, 0, +1, +2, +3$

Hence, the answer is 3

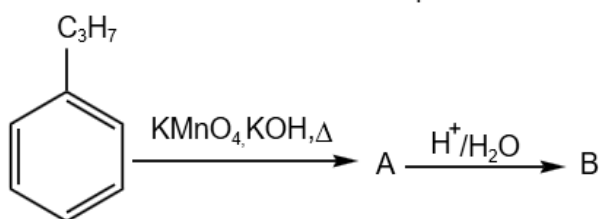
Q.33. One coulomb charge is passed through $AgNO_3$ solution during electrolysis. Find mass of silver (in mg) deposited at the electrode. Give answer in nearest integer.

Answer: 1

Solution: $m = ZIt$
 $q = It$
 $\frac{\text{equivalent mass}}{\text{Faraday}} = \frac{\text{mass}}{n_{\text{factor}} \times 96500}$
 $= \frac{108}{96500} \times 10^{-3} \text{ g}$
 $= \frac{108}{96.5} \text{ g} = 1.11 \text{ g}$

Hence, the answer is 1

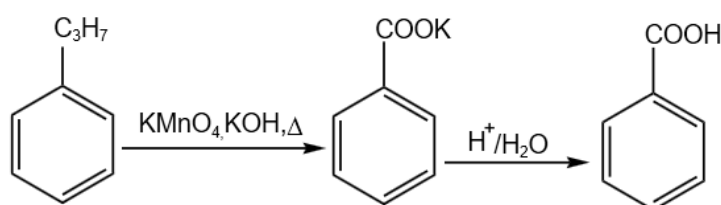
Q.34. Consider the following reaction



The number of pi bonds in product B is

Answer: 4

Solution: The propyl benzene on reaction with $KMnO_4 / KOH$ give potassium benzoate. Potassium benzoate in acidic medium results benzoic acid.



The number of pi bonds in benzoic acid is 4.

Q.35. How many can give H_2 gas from dilute acid of Ti^{2+} , Cr^{2+} , V^{2+} , Mn^{2+} , Fe^{2+} , Co^{2+}

Answer: 5

Solution: $Mn^{2+}: 3d^5$

It is in a stable state. Hence, it will not willing to liberate hydrogen gas.

Ti^{2+} , Cr^{2+} , V^{2+} , Fe^{2+} , Co^{2+} will be able to liberate hydrogen gas.

Hence, the answer is 5



Q.36. Find the coefficient of x^0 in the expansion of $\left(\frac{\frac{1}{35}}{x} + \frac{x}{\frac{1}{53}}\right)^{12}$

A) ${}^{12}C_7 \left(\frac{3^6/5}{5^2}\right)$ B) ${}^{12}C_6 \left(\frac{3^6/5}{5^2}\right)$

C) ${}^{12}C_6 \left(\frac{3}{5^2}\right)$ D) ${}^{12}C_5 \left(\frac{3^2}{5^2}\right)$

Answer: ${}^{12}C_6 \left(\frac{3^6/5}{5^2}\right)$

Solution: Given,

Binomial expression $\left(\frac{\frac{1}{35}}{x} + \frac{x}{\frac{1}{53}}\right)^{12}$

Now, $T_{r+1} = {}^{12}C_r \left(\frac{\frac{1}{35}}{x}\right)^{12-r} \left(\frac{x}{\frac{1}{53}}\right)^r$

$\Rightarrow T_{r+1} = {}^{12}C_r \left(\frac{1}{35}\right)^{12-r} \left(\frac{-1}{5 \cdot 3}\right)^r x^{2r-12}$

Now, for constant term we get, $r = 6$

So, $T_7 = {}^{12}C_6 \left(3^{1/5}\right)^{12-6} \left(5^{-1/3}\right)^6$

$\Rightarrow T_7 = {}^{12}C_6 \left(\frac{3^6/5}{5^2}\right)$

Q.37. Find the minimum value of k so that $4^{x+1} + 4^{1-x}, \frac{k}{2}, 16^x + 16^{-x}$ are in AP.

A) 5 B) 16

C) 10 D) 12

Answer: 10

Solution: Given: $4^{x+1} + 4^{1-x}, \frac{k}{2}, 16^x + 16^{-x}$ are in AP.

$\Rightarrow 4^{x+1} + 4^{1-x} + 16^x + 16^{-x} = \frac{k}{2} \times 2$

$\Rightarrow k = 4 \left(4^x + \frac{1}{4^x}\right) + \left(4^{2x} + \frac{1}{4^{2x}}\right)$

We know that, $AM \geq GM$

$\Rightarrow \frac{4^x + \frac{1}{4^x}}{2} \geq 4^x \times \frac{1}{4^x}, \frac{4^{2x} + \frac{1}{4^{2x}}}{2} \geq 4^{2x} \times \frac{1}{4^{2x}}$

$\Rightarrow 4^x + \frac{1}{4^x} \geq 2, 4^{2x} + \frac{1}{4^{2x}} \geq 2$

$\Rightarrow k = 4(2) + (2)$

$\Rightarrow k \geq 10$

Hence, the minimum value of k is 10.

Q.38. If $x|x| + 2|x - 5| + 2x + 3 = 0$. Find the value of x .

A) $-\sqrt{13}$

B) $\sqrt{13}$



C) 5

D) 12

Answer: $-\sqrt{13}$

Solution: Given: $x|x| + 2|x - 5| + 2x + 3 = 0$

Case-I: When $x \leq 0$

$$\Rightarrow -x^2 + 10 - 2x + 2x + 3 = 0$$

$$\Rightarrow x^2 = 13$$

$$\Rightarrow x = -\sqrt{13}$$

Case-II: When $0 < x < 5$

$$\Rightarrow x^2 + 10 - 2x + 2x + 3 = 0$$

$$\Rightarrow x^2 = -13$$

So, no solution in this case.

Case-III: When $x \geq 5$

$$\Rightarrow x^2 + 2x - 10 + 2x + 3 = 0$$

$$\Rightarrow x^2 + 4x - 7 = 0$$

$$\Rightarrow x = \frac{-4 \pm \sqrt{16 + 28}}{2}$$

$$\Rightarrow x = \frac{-4 \pm 2\sqrt{11}}{2}$$

So, no solution in this case.

Hence, $x = -\sqrt{13}$

Q.39. The 50th word in the dictionary using the letters *B, B, H, J, O*

A) *JHBBO*

B) *OBBJH*

C) *BBJOH*

D) *OBBHJ*

Answer: *OBBJH*

Solution: Given,

Letters *B, B, H, J, O*

Now, word starting with *B* we have,

$$B _ _ _ _ \rightarrow 4! = 24$$

Similarly, with *H* we get,

$$H _ _ _ _ \rightarrow \frac{4!}{2} = 12 \text{ \{as 2 B are repeating\}}$$

And with *J* we get,

$$J _ _ _ _ \rightarrow \frac{4!}{2} = 12 \text{ \{as 2 B are repeating\}}$$

So, 48 words have already come, now 49th and 50th word will be *OBBHJ* & *OBBJH* respectively,

Hence, *OBBJH* is required answer.

Q.40. Find the area bounded by $y = -2|x|$ and $y = x|x|$

A) $\frac{1}{2}$

B) $\frac{2}{3}$

C) $\frac{4}{3}$

D) $\frac{1}{3}$

Answer: $\frac{4}{3}$



Solution: Given: $y = -2|x|$ and $y = x|x|$

$$\Rightarrow -2|x| = x|x|$$

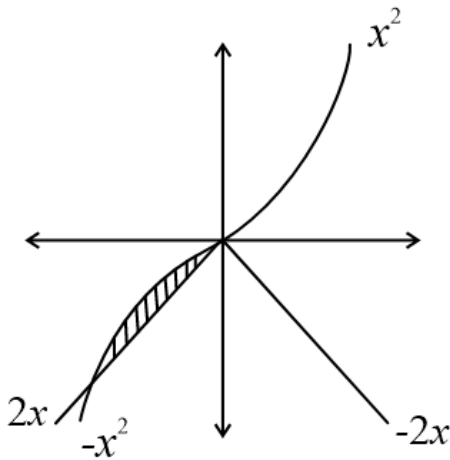
$$\Rightarrow |x|(x+2) = 0$$

$$\Rightarrow x = 0, -2$$

$$\Rightarrow y = 0, -4$$

So, the points of intersection are $(0, 0)$ and $(-2, -4)$.

$$\Rightarrow y = \begin{cases} -2x & x \geq 0 \\ 2x & x < 0 \end{cases} \text{ and } y = \begin{cases} x^2 & x \geq 0 \\ -x^2 & x < 0 \end{cases}$$



So, the required area is given by,

$$A = \left| \int_{-2}^0 (-x^2 - (2x)) dx \right|$$

$$\Rightarrow A = \left| \left[\frac{-x^3}{3} - x^2 \right]_{-2}^0 \right|$$

$$\Rightarrow A = \left| \frac{8}{3} - 4 \right|$$

$$\Rightarrow A = \frac{4}{3} \text{ square units.}$$

Q.41. $ABCD$ and $A EFG$ are two squares of side length 4 and 2 respectively. If E lies on line segment AB and F lies on diagonal AC , find radius of circle passing through F and touching BC and CD

A) $4 + 2\sqrt{2}$

B) $4 - 2\sqrt{2}$

C) $2\sqrt{2}$

D) $\sqrt{2}$

Answer: $4 - 2\sqrt{2}$



Solution: Given,

$$\text{Equation } P(x) = ax^2 + bx + c = 0,$$

Now, for real and distinct roots $b^2 > 4ac$ where $a, b, c \in A = \{1, 2, 3, 4, 5, 6\}$

Now, for $b = 1$ & 2 there will be no cases,

For $b = 3$ we get, $ac = \{(1, 1), (1, 2), (2, 1)\} \rightarrow 3$ ways

For $b = 4$ we get, $ac = \{(1, 1), (1, 2), (2, 1), (1, 3), (3, 1)\} \rightarrow 5$ ways

For $b = 5$ we get,

$ac = \{(1, 1), (1, 2), (2, 1), (1, 3), (3, 1), (2, 2), (1, 4), (1, 4), (2, 3), (3, 2), (5, 1), (1, 5), (1, 6), (6, 1)\} \rightarrow 14$ ways

And for $b = 6$ we get,

$ac = \{(1, 1), (1, 2), (2, 1), (1, 3), (3, 1), (2, 2), (1, 4), (1, 4), (2, 3), (3, 2), (5, 1), (1, 5), (1, 6), (6, 1), (2, 4), (4, 2)\} \rightarrow 16$ ways

So, total possible ways are $3 + 5 + 14 + 16 = 38$

Hence, probability is given by $\frac{38}{6^3} = \frac{19}{108}$

Q.46.

If $2x^2 - x + 2 = 0$ and its one root is a , then $\lim_{x \rightarrow \frac{1}{a}} \frac{16(1 - \cos(2x^2 - x + 2))}{(ax - 1)^2}$ is equal to:

A) $\frac{20(1-a^2)^2}{a^3}$

B) $\frac{8(1-a^2)^2}{a^3}$

C) $\frac{32(1-a^2)^2}{a^4}$

D) $\frac{16(1-a^2)^2}{a^4}$

Answer: $\frac{32(1-a^2)^2}{a^4}$



Answer: 14

Solution: Given,

$$\text{Equation of line } \frac{x-1}{2} = \frac{y+1}{3} = \frac{z-2}{5}$$

Now, any point on the above line is given by,

$$Q(2r+1, 3r-1, 5r+2)$$

Now, the direction ratio of line from point $P(8, 5, 7)$ to point $Q(2r+1, 3r-1, 5r+2)$ is given by $(2r-7, 3r-6, 5r-5)$

Now, using the perpendicular condition as PQ is perpendicular to the given line we get,

$$2(2r-7) + 3(3r-6) + 5(5r-5) = 0$$

$$\Rightarrow 38r = 57$$

$$\Rightarrow r = \frac{3}{2}$$

$$\text{So, point } Q(2r+1, 3r-1, 5r+2) = Q\left(4, \frac{7}{2}, \frac{19}{2}\right)$$

Now, using the midpoint formula to find the image $P'(\alpha, \beta, \gamma)$ we get,

$$\frac{\alpha+8}{2} = 4, \frac{\beta+5}{2} = \frac{7}{2}, \frac{\gamma+7}{2} = \frac{19}{2}$$

$$\Rightarrow \alpha = 0, \beta = 2, \gamma = 12$$

$$\Rightarrow \alpha + \beta + \gamma = 14$$

Q.49. If line L is perpendicular to $y = 2x + 10$ such that it touches the parabola $y^2 = 4(x - 9)$ then find the square of distance between the point of contact and origin

Answer: 185

Solution: Given,

A line L is perpendicular to $y = 2x + 10$ such that it touches the parabola $y^2 = 4(x - 9)$,

So, let the line $2y + x = c$ be line perpendicular to the given line,

Now, given $2y + x = c$ touches the parabola $y^2 = 4(x - 9)$

$$\text{So, } \left(\frac{c-x}{2}\right)^2 = 4(x-9)$$

$$\Rightarrow x^2 - 2(c+8)x + c^2 + 144 = 0$$

Making discriminant $D = 0$ we get,

$$\Rightarrow 4(c+8)^2 - 4(c^2 + 144) = 0$$

$$\Rightarrow c = 5$$

So, the equation of line will be $2y + x = 5$ and point of contact will be $(13, -4)$

Hence, distance of point of point $(13, -4)$ from origin is $\sqrt{169 + 16} = \sqrt{185}$.

Q.50. If $|\vec{a}| = 2$, $|\vec{b}| = 3$ & $\vec{a} = \vec{b} \times \vec{c}$, then find the minimum value of $9|\vec{c} - \vec{a}|^2$

Answer: 40



Solution: Given,

$$|\vec{a}| = 2, |\vec{b}| = 3,$$

$$\text{And } \vec{a} = \vec{b} \times \vec{c}$$

$$\Rightarrow \vec{a} \cdot \vec{b} = 0 \text{ \& } \vec{a} \cdot \vec{c} = 0$$

$$\text{Now, solving } |\vec{c} - \vec{a}|^2 = |\vec{c}|^2 + |\vec{a}|^2 + 2\vec{c} \cdot \vec{a}$$

$$\Rightarrow |\vec{c} - \vec{a}|^2 = |\vec{c}|^2 + 4 \dots \dots \dots (i)$$

$$\text{Now, solving } |\vec{a}| = |\vec{b} \times \vec{c}|$$

$$\Rightarrow |\vec{a}| = |\vec{b}| |\vec{c}| \sin \theta$$

$$\Rightarrow |\vec{c}| = \frac{2}{3 \sin \theta}$$

$$\Rightarrow |\vec{c}|^2 = \frac{4}{9 \sin^2 \theta} \dots \dots \dots (ii)$$

From equation (i) & (ii) we get,

$$|\vec{c} - \vec{a}|^2 = \frac{4}{9 \sin^2 \theta} + 4$$

$$\Rightarrow |\vec{c} - \vec{a}|^2 \geq \frac{4}{9} + 4$$

$$\Rightarrow |\vec{c} - \vec{a}|^2 \geq \frac{40}{9}$$

$$\Rightarrow 9|\vec{c} - \vec{a}|^2 \geq 40$$